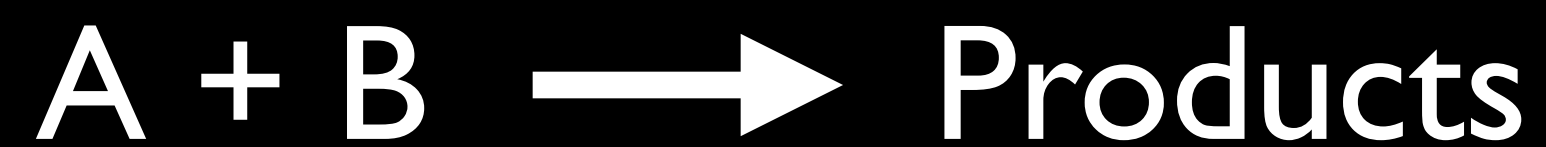
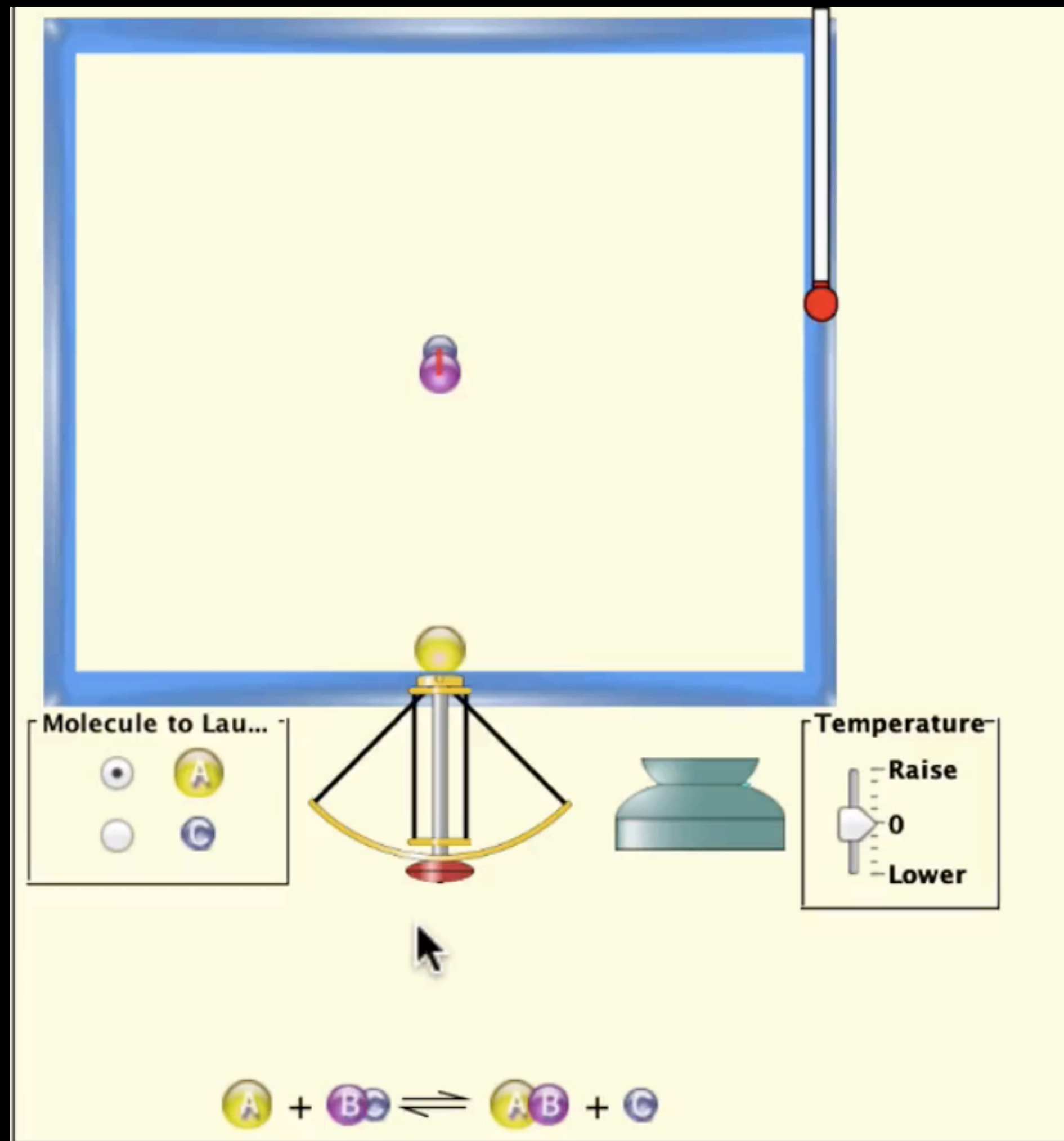
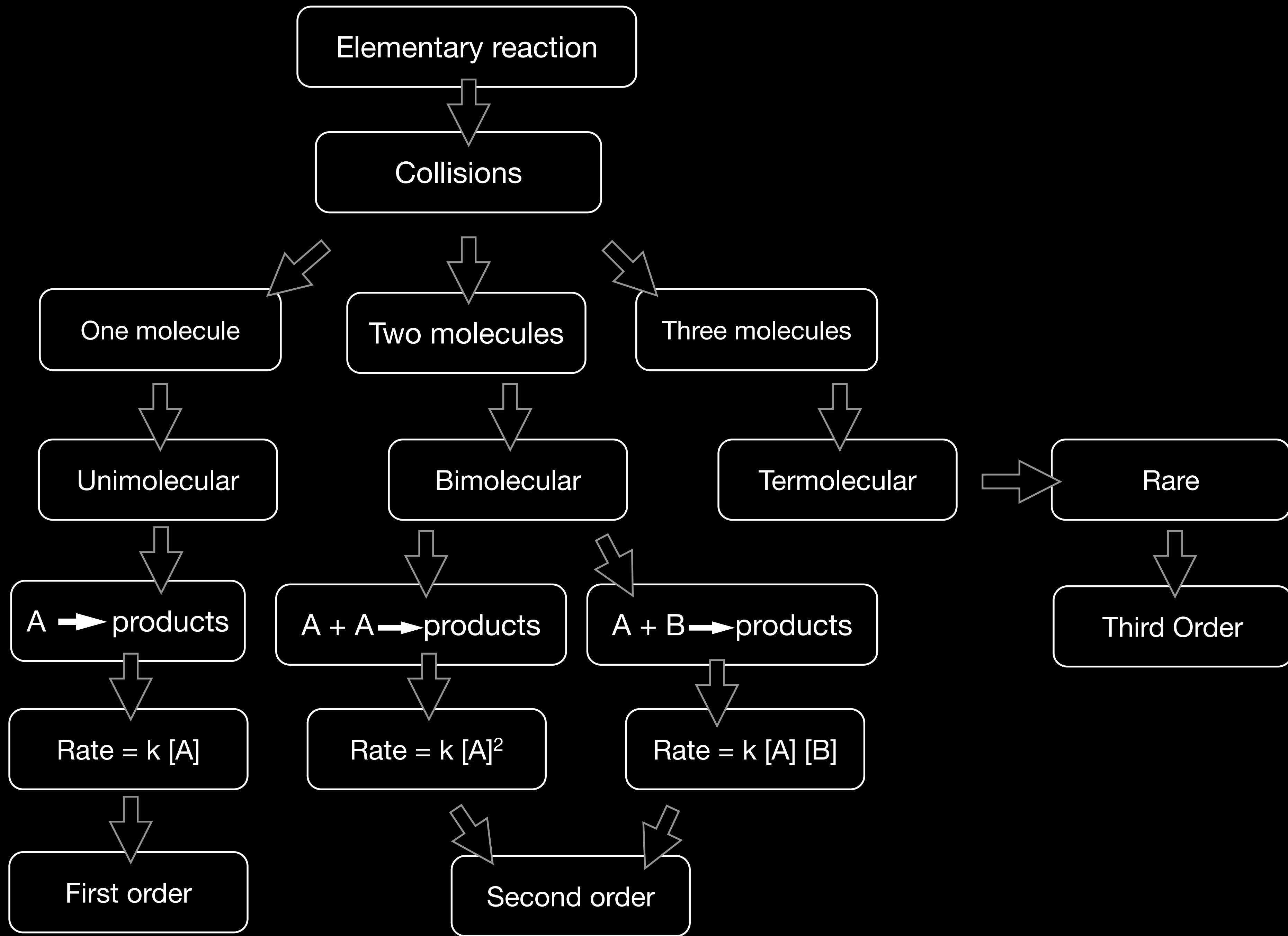


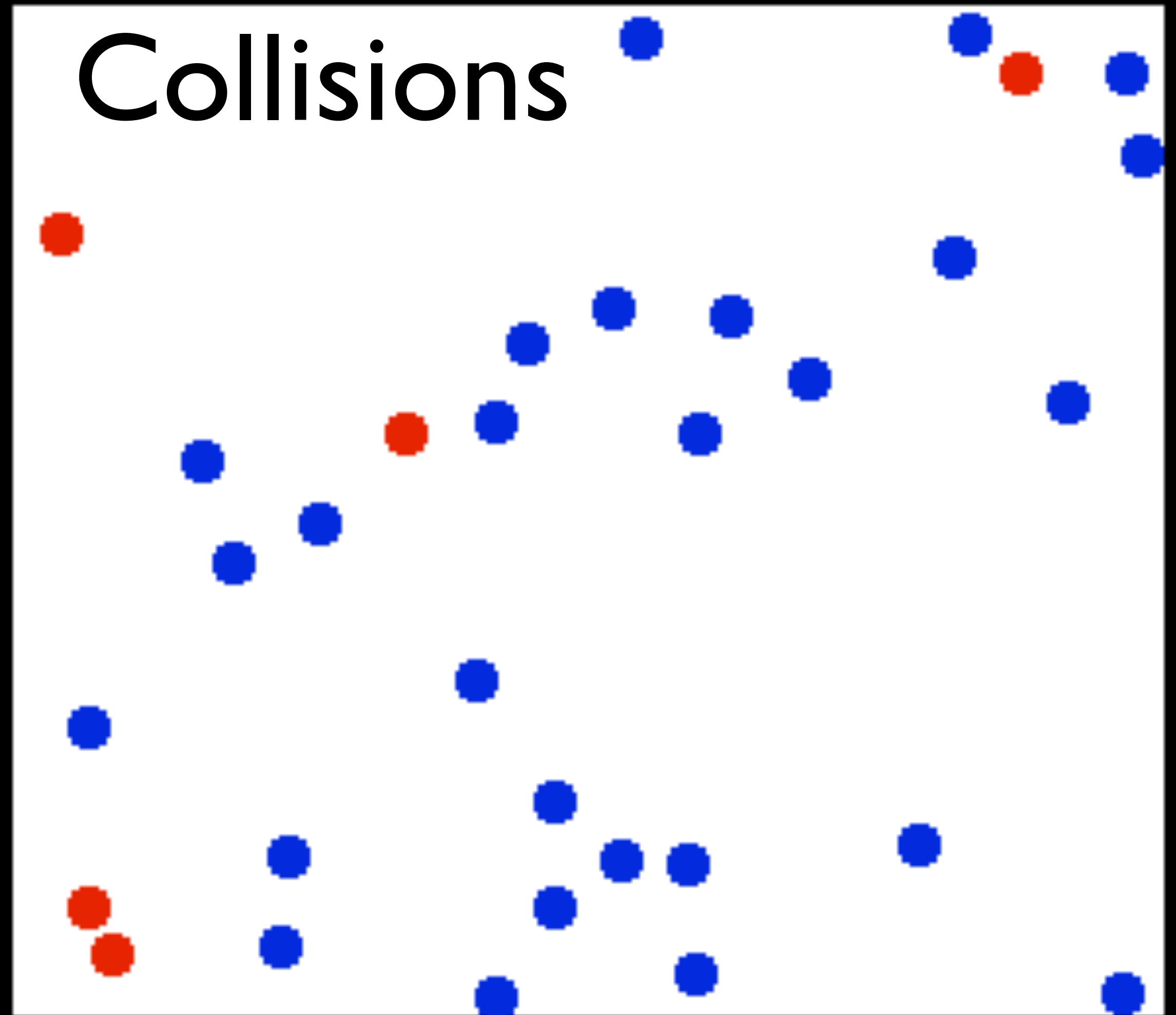
# Elementary Reactions

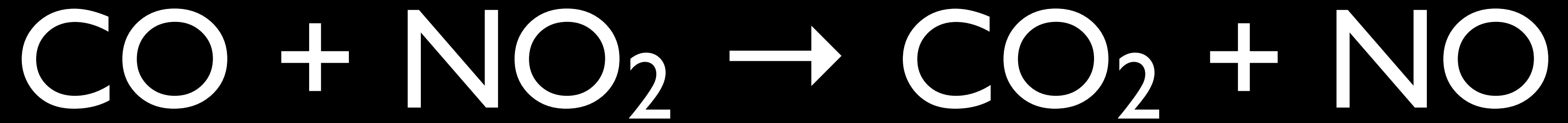


Chemistry Essentials - 038



- Energy
- Orientation





Stoichiometry

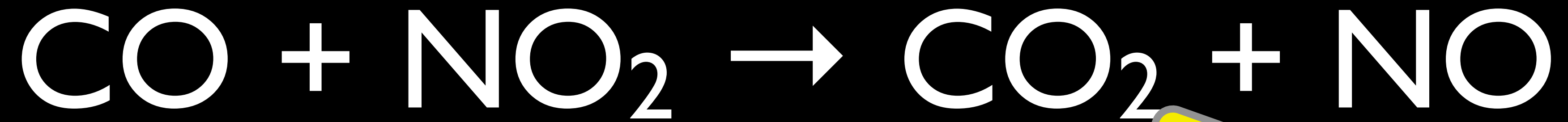


Predictable

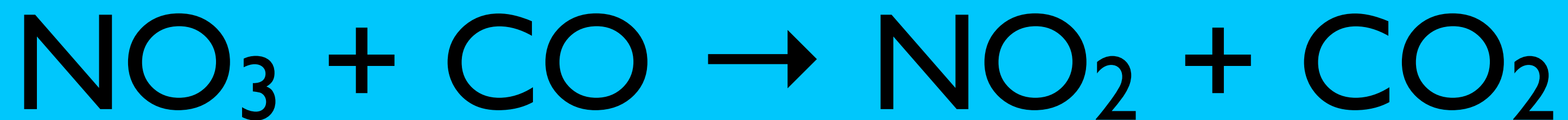
Rate Law



Measureable



Elementary Reaction

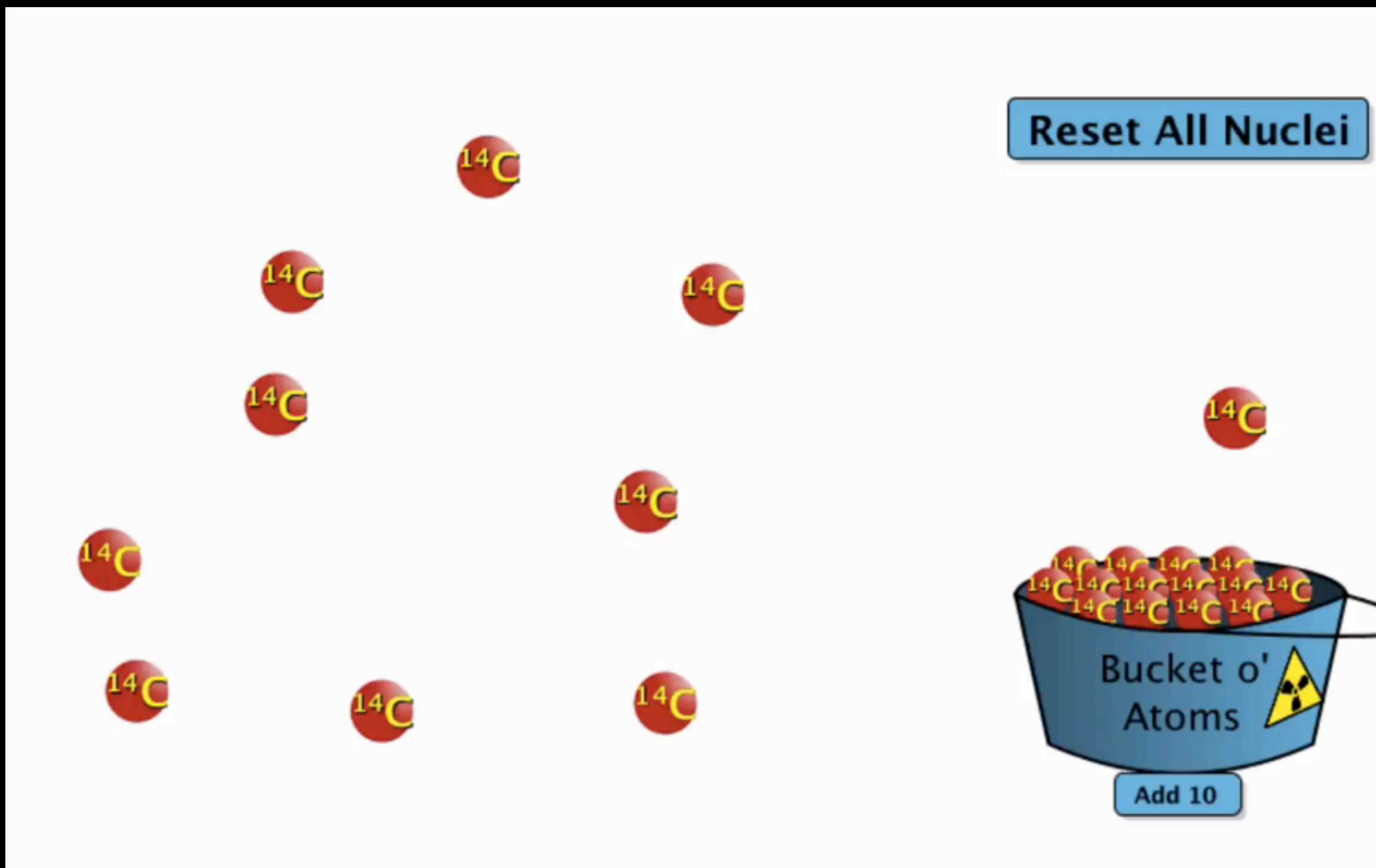


Reaction Mechanism

Predictable

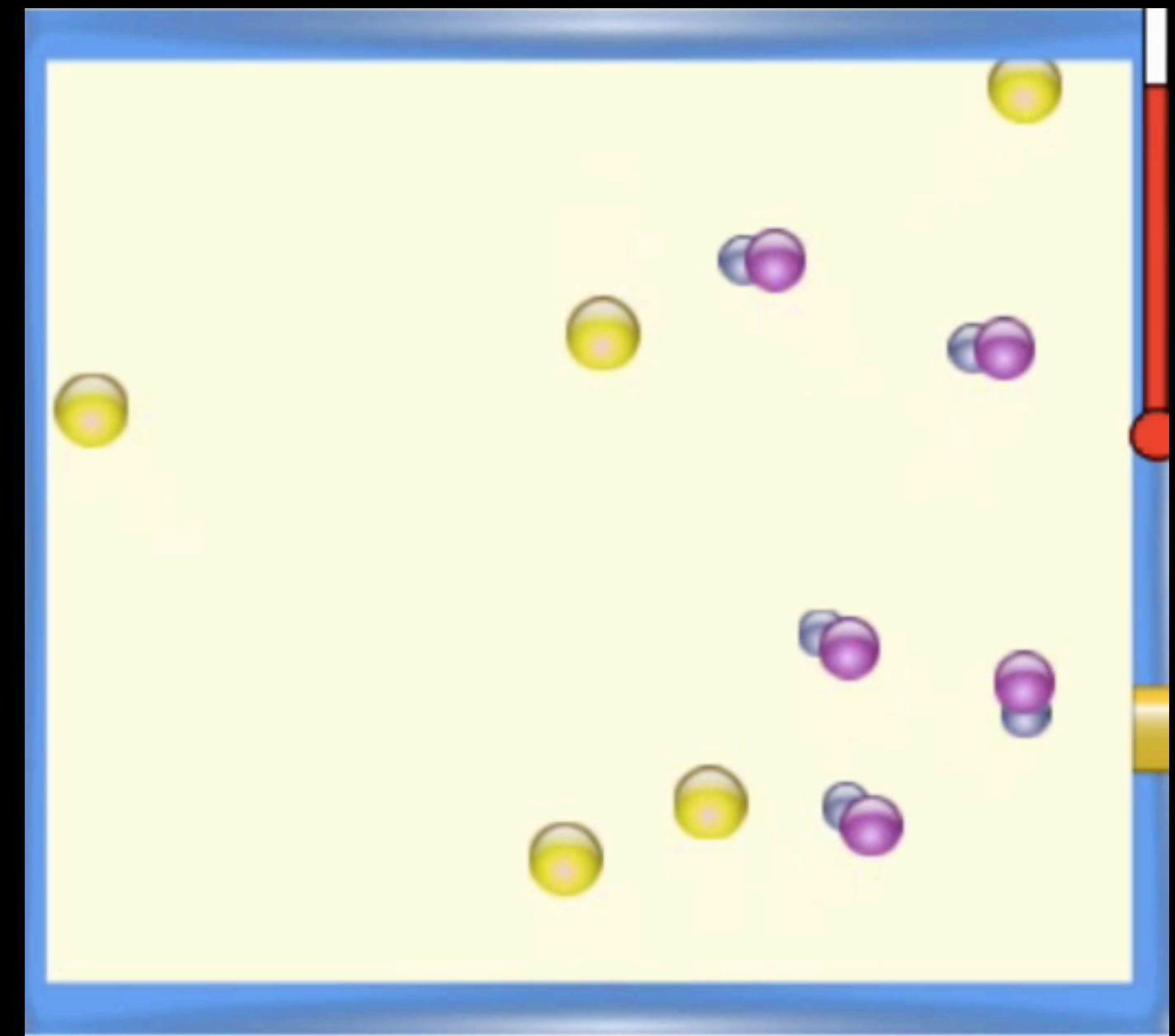
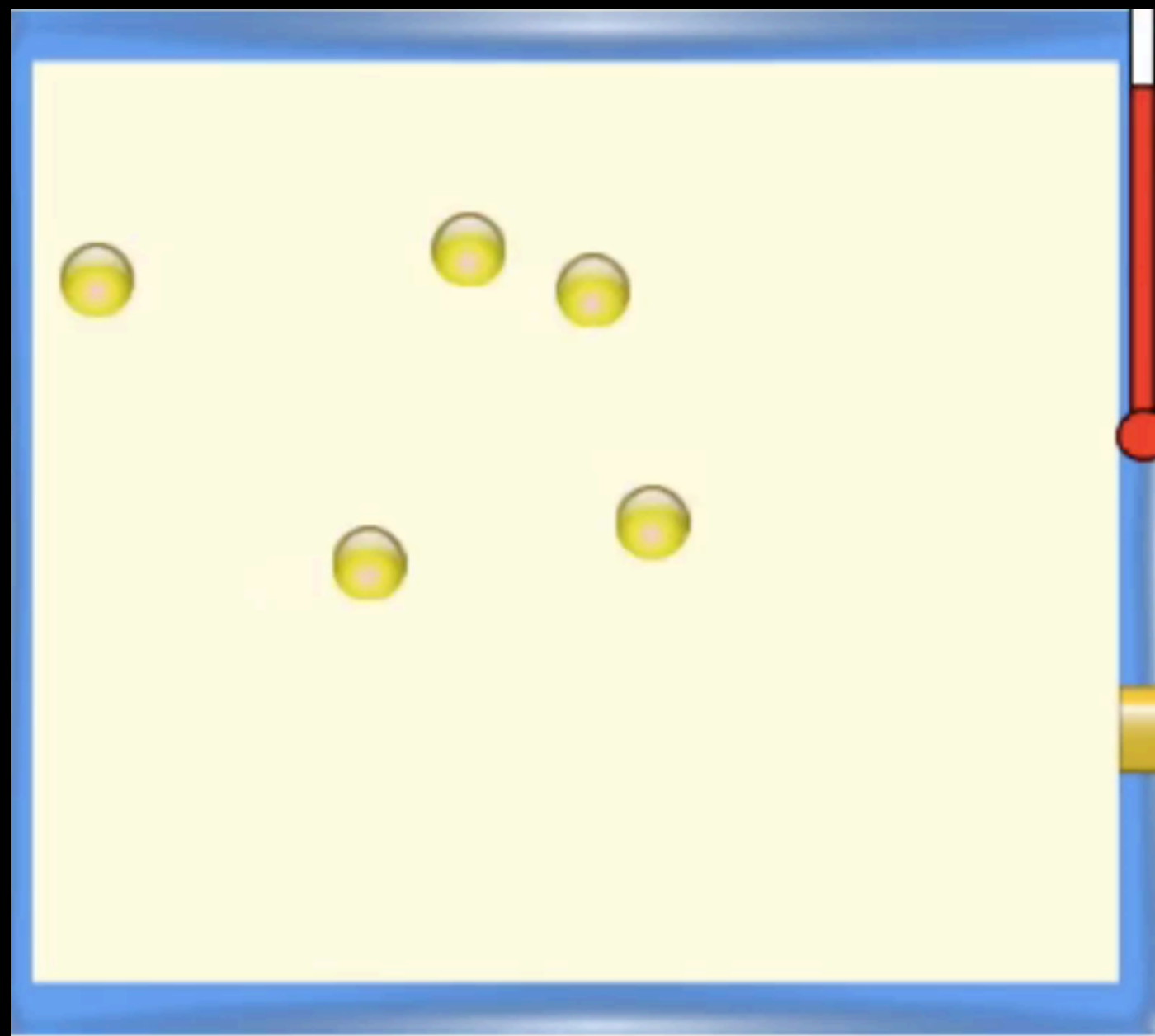
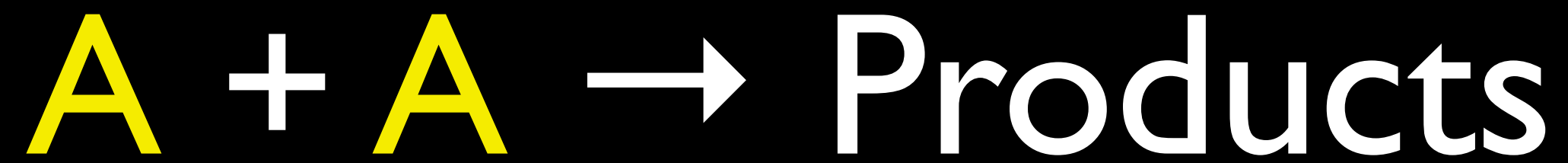
Measureable

# Unimolecular Elementary Reaction



$$\text{Rate} = k [A]$$

# Bimolecular Elementary Reaction



$$\text{Rate} = k [A]^2$$

$$\text{Rate} = k [A][B]$$

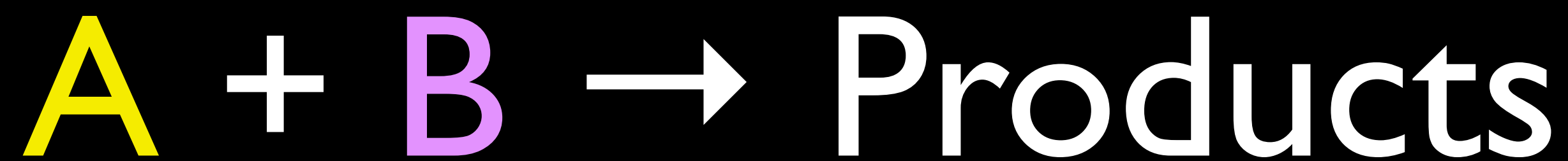
# Termolecular Elementary Reaction



$$\text{Rate} = k [A][B][C]$$



Did you learn?



$$\text{Rate} = k [A][B]$$

To connect the rate law for an elementary reaction to the frequency and success of molecular collisions.

## Acknowledgements

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