

Precipitation 2

- Worksheet -

1. Write ionic equations for the following reactions.

- (1). Silver nitrate and sodium chloride produce silver chloride precipitation.
- (2). Lead(II) nitrate and potassium iodide produce lead(II) iodide precipitation.
- (3). Barium chloride and sodium sulfate produce barium sulfate precipitation.
- (4). Silver nitrate and potassium chromate produce silver chromate precipitation.
- (5). Calcium chloride and sodium carbonate produce calcium carbonate precipitation.
- (6). Chromium(III) chloride and sodium hydroxide produce chromium(III) hydroxide precipitation.

2. Watch the film, and try to find the colors of different ions and precipitate.

Ion	Color	Precipitate	Color
Silver		Silver chloride	
		Silver chromate	
Lead (II)		Lead iodide	
Barium		Barium sulfate	
Calcium		Calcium carbonate	
Chromium (III)		Chromium hydroxide	

3. Can you deduce any solubility rules according to the table above? Please check "your rules" with the solubility chart on your text book.

4. Determine whether the two substances will produce a precipitate.

